Related to chapter III

1) At constant temperature, in spontaneous phenomena the entropy change is increasing or decreasing

----increasing

2) Which is the discipline that bridges microscopic and macrospic descriptions

--thermodinamics

3) Why thermodynamics is so important in biology

---explain function and the ability to interact

4) The field of application/s of thermodynamics

---everything

5) The definition of an equilibrium constant with constraints

---DeltaG=-RTlnKeq

6) The role of the Gibbs free energy change

Is the amount of potential energy used in the trasformation

7) What is a standard condition

--273,15 k 1 bar

8) How are entropy, enthalpy and Gibbs free energy related

U(internal energy)=Q-W

H(enthalpy)=U+pv

DeltaS(etropiy)=DeltaQ(reversible)\T

G(gibbs)=H-TS

9) A definition of pH

Ph=-log[H+]

10) A definition of hydrophobicity

**A plane definition of the hydrophobic effect**

The hydrophobic effect is the observed tendency of nonpolar substances to aggregate in aqueous solution

and exclude water mole

cules.

This occurs because interactions between the hydrophobic molecules enable

the displaced water molecules to make hydrogen bonds more freely with each other, thereby decreasing

the overall free energy.

11) A definition of “hydrophobic effect”

**A plane definition of the hydrophobic effect**

The hydrophobic effect is the observed tendency of nonpolar substances to aggregate in aqueous solution

and exclude water mole

cules.

This occurs because interactions between the hydrophobic molecules enable

the displaced water molecules to make hydrogen bonds more freely with each other, thereby decreasing

the overall free energy.